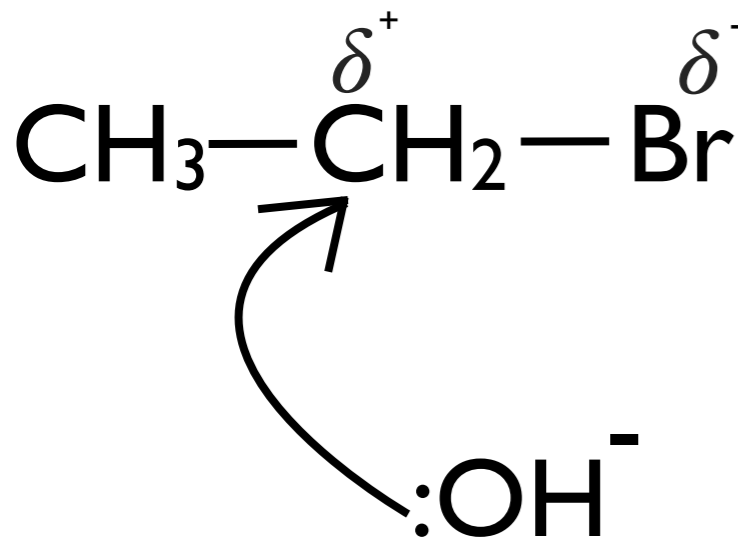
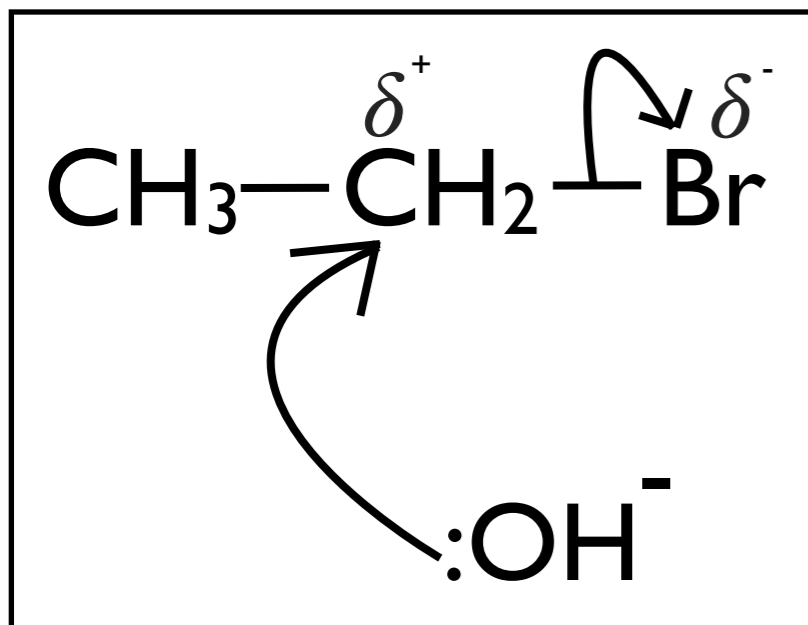


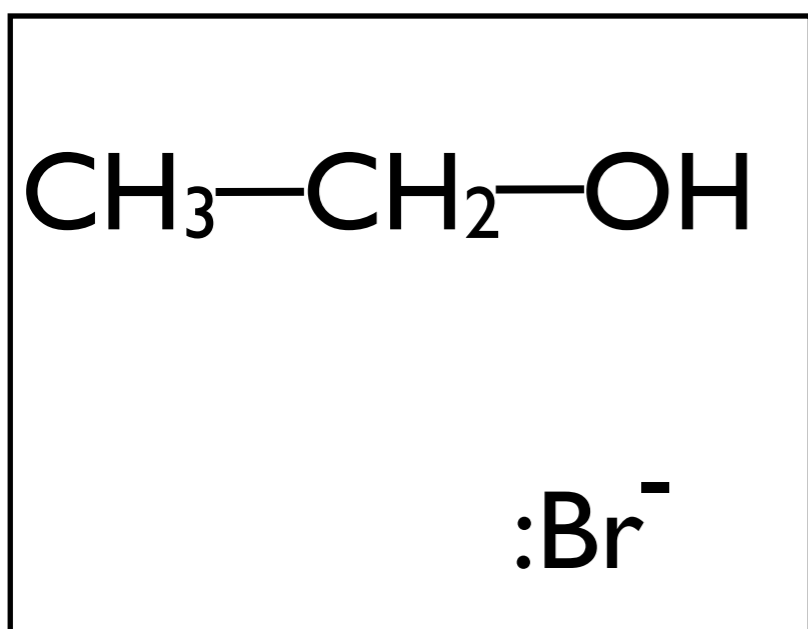
Step 1: The negatively charged hydroxide nucleophile is attracted to the positive dipole on the middle carbon atom. This dipole exists because of the difference in electronegativity between carbon and bromine. You must show the lone pair on the nucleophile and the delta symbols which indicate the polarity of the dipole.



Step 2: The lone pair of the hydroxide nucleophile forms a covalent bond with the carbon atom with the positive dipole. This is indicated by the direction of the curly arrow.



Step 3: The covalent bond between the carbon atom and the bromine atom breaks. Both electrons are attracted to the bromine atom.



Step 4: The hydroxide nucleophile has substituted for the bromine atom, which is now a bromine ion as it gained both electrons used to form its covalent bond with carbon.